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Published in:
The Journal of Physical Chemistry Letters

Link to article, DOI:
[10.1021/acs.jpcllett.6b00091](https://doi.org/10.1021/acs.jpcllett.6b00091)

Publication date:
2016

Document Version
Peer reviewed version

[Link back to DTU Orbit](#)

Citation (APA):
Kumar, N., & Siegel, D. J. (2016). Interface-Induced Renormalization of Electrolyte Energy Levels in Magnesium Batteries. *The Journal of Physical Chemistry Letters*, 7, 874-881. <https://doi.org/10.1021/acs.jpcllett.6b00091>

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Interface-Induced Renormalization of Electrolyte Energy Levels in Magnesium Batteries

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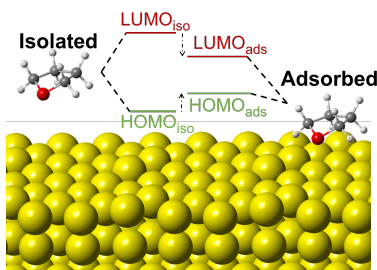
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Abstract: A promising strategy for increasing the energy density of Li-ion batteries is to substitute a multivalent (MV) metal for the commonly-used lithiated carbon anode. Magnesium is a prime candidate for such a MV battery due to its high volumetric capacity, abundance, and limited tendency to form dendrites. One challenge that is slowing the implementation of Mg-based batteries, however, is the development of efficient and stable electrolytes. Computational screening for molecular species having sufficiently wide electrochemical windows is a starting point for the identification of optimal electrolytes. Nevertheless, this window can be altered via interfacial interactions with electrodes. These interactions are typically omitted in screening studies, yet they have the potential to generate large shifts to the HOMO and LUMO of the electrolyte components. The present study quantifies the stability of several common electrolyte solvents on model electrodes of relevance for Mg batteries. Many-body perturbation theory calculations based on the G_0W_0 method were used to predict shifts in a solvent's electronic levels arising from interfacial interactions. In molecules exhibiting large dipole moments our calculations indicate that these interactions reduce the HOMO-LUMO gap by $\sim 25\%$ (compared to isolated molecules). We conclude that electrode interactions can narrow an electrolyte's electrochemical window significantly, thereby accelerating redox decomposition reactions. Accounting for these interactions in screening studies presents an opportunity to refine predictions of electrolyte stability.

TOC Figure:



Electrolytes are essential components of batteries.¹ Modern batteries based on the Li-ion chemistry employ electrolytes that allow for facile ionic transport of Li⁺ between the negative and positive electrodes, while acting as an electronic insulator that prevents short-circuiting. Common electrolytes are comprised of Li salts such as LiPF₆ and organic solvents such as ethylene carbonate and dimethyl carbonate.

In addition to having appropriate transport properties, the electronic structure of the electrolyte should be compatible with the electrochemical window of the cell.² Figure 1 illustrates the relationship between electron energy levels in electrodes and in the electrolyte. Here, the redox levels (i.e. Fermi energies) of the negative electrode (anode) and positive electrode (cathode) are represented by their electrochemical potentials, μ_A and μ_C , respectively. Electrochemical stability requires that two criteria are satisfied: first, the energy gap of the electrolyte, E_g , between the highest occupied molecular orbital (HOMO) and lowest unoccupied molecular orbital (LUMO) should be larger than the voltage difference, V_{OC} , between electrodes: $E_g > V_{OC}$. Second, the HOMO and LUMO levels should be positioned ‘outside’ of the electrochemical potentials of the electrodes: $E_{LUMO} > \mu_A$ and $E_{HOMO} < \mu_C$. A LUMO that is ‘too low’ will result in reduction of the electrolyte by the anode; similarly, a HOMO that is too high will allow oxidation of the electrolyte by the positive electrode. This oxidation and/or reduction is generally undesirable, as it can lead to irreversible bond breaking that degrades and/or dries out the electrolyte, leading to cell failure.

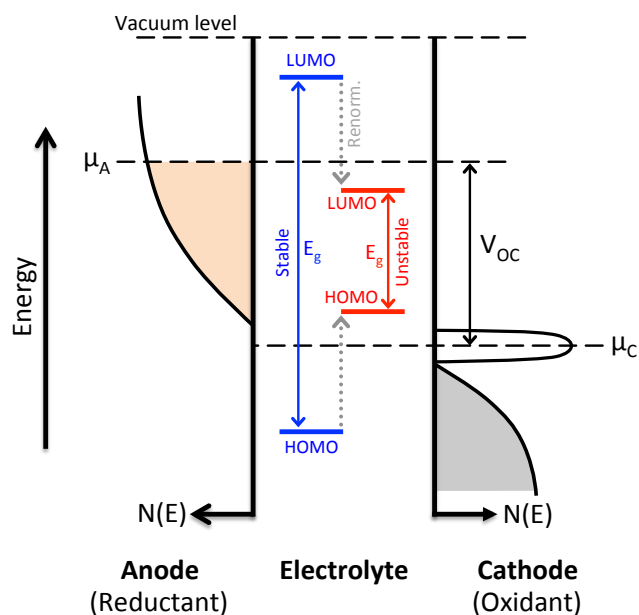


Figure 1. Position of electrolyte HOMO/LUMO levels (center) with respect to the density of states of a battery anode (left) and cathode (right). The Fermi level of the anode (assumed to be a metal) is denoted by μ_A ; μ_C represents the redox level of the cathode. V_{OC} represents the open circuit voltage. An electrolyte that is stable against oxidation by the cathode and reduction by the anode is illustrated in blue. Gray arrows depict the impact of interfacial interactions between the electrolyte molecules and electrode surfaces, which renormalize the HOMO/LUMO levels. Red levels illustrate an unstable electrolyte. E_g is the HOMO/LUMO energy gap. Adapted from Ref. 2.

Due to the highly electropositive nature of lithium, many electrolytes employed in Li-ion batteries are not stable with respect to reduction by the negative electrode³⁻⁴. The fact that these batteries function very well, however, can be traced to the formation of a protective solid-electrolyte interphase, SEI, during the initial cycling of the battery.⁵ The SEI “kinetically stabilizes” the electrode/electrolyte interface by inhibiting further reductive decomposition of the electrolyte, while allowing for Li ion transport.

Although Li-ion batteries represent the state-of-the-art in electrical energy storage, next-generation applications such as extended-range battery electric vehicles would strongly benefit from batteries having higher energy densities.⁶⁻⁷ A promising alternative to the Li-ion chemistry would employ a negative electrode whose active material can adopt a multivalent (MV) oxidation state, such as 2+ or 3+. In this regard batteries based on divalent magnesium (Mg) have received increasing attention.^{6, 8-15} The theoretical volumetric capacity of a Mg anode, 3833 mAh/cm³, is nearly double that of monovalent Li, 2046 mAh/cm³.^{8, 16-17} A Mg-based battery may also possess advantages in safety and cost compared to Li systems. Regarding cost, magnesium is the 5th most abundant element in the earth's crust, whereas Li is ranked 22nd. Regarding safety, anodes based on Mg metal appear to be less susceptible to dendrite formation than those based on Li.⁹

Despite these benefits, several aspects of the Mg-ion system present challenges to its realization as a viable secondary battery. For example, few cathode materials are capable of intercalating Mg ions at high capacities, fast rates, and at sufficiently positive voltages.¹⁰ The stability of the anode/electrolyte interface also remains a point of concern: In contrast to Li-ion systems, where the formation of an SEI limits the rate of electrode corrosion and electrolyte decomposition, the existence and desirability of an SEI on the surface of a Mg anode remains a matter of debate.^{11-12, 18-20} Aurbach has argued that an SEI is undesirable in Mg batteries since the resulting film is not amenable to facile Mg-ion transport.⁸

Figure 1 and the preceding discussion highlight the importance of the electrolyte's electronic structure (i.e., HOMO-LUMO positions) in controlling electrolyte decomposition and the tendency for SEI formation. Modern electronic structure methods can calculate these redox levels with a high degree of accuracy at relatively modest computational cost. Consequently,

computational screening is being increasingly employed to identify suitable electrolyte components.²¹⁻³⁰ Most studies of this type aim to predict the HOMO/LUMO levels of a particular electrolyte component, for example, by screening individual solvent or salt molecules. While evaluating the properties of isolated molecules is a reasonable starting point, interactions between electrolyte molecules (e.g., solvent-salt²¹) and electrode surfaces can alter HOMO/LUMO levels substantially.³¹⁻³³ The inclusion of these interactions into screening analyses will lead to more accurate predictions of the stability window of candidate electrolytes.³⁴ Towards this goal, the present study aims to quantify the renormalization of HOMO/LUMO levels of several common electrolyte solvents that arise from interactions with prototypical Mg battery electrodes. As illustrated in Figure 1, electrode interactions are critical for electrolyte stability as oxidation and/or reduction of the electrolyte occurs through electron transfer at interfaces with solid electrodes.

It is well known that the positions of HOMO/LUMO levels, and resulting gap between them, are poorly described by semi-local density functionals.³¹⁻³² Although hybrid functionals yield some improvement in the estimation of these quantities for isolated molecules, they fail to capture correlation effects such as image charge interactions at electrode/molecule interfaces. Garcia-Lastra *et al.*³¹⁻³² demonstrated the limitations of semi-local and hybrid functionals in determining the renormalization of HOMO/LUMO levels of a benzene molecule adsorbed on several metallic and semiconducting surfaces. It was shown that an accurate determination of the frontier orbitals required the use of quasi-particle methods based on many-body perturbation theory. Related work by Neaton *et al.*³³ demonstrated that the bandgap of benzene adsorbed on graphite (0001) was essentially unchanged with respect to the gas-phase molecule within the

local density approximation. This behavior differed significantly from the predictions of GW calculations, which yielded a significant (30%) narrowing of the HOMO/LUMO gap.

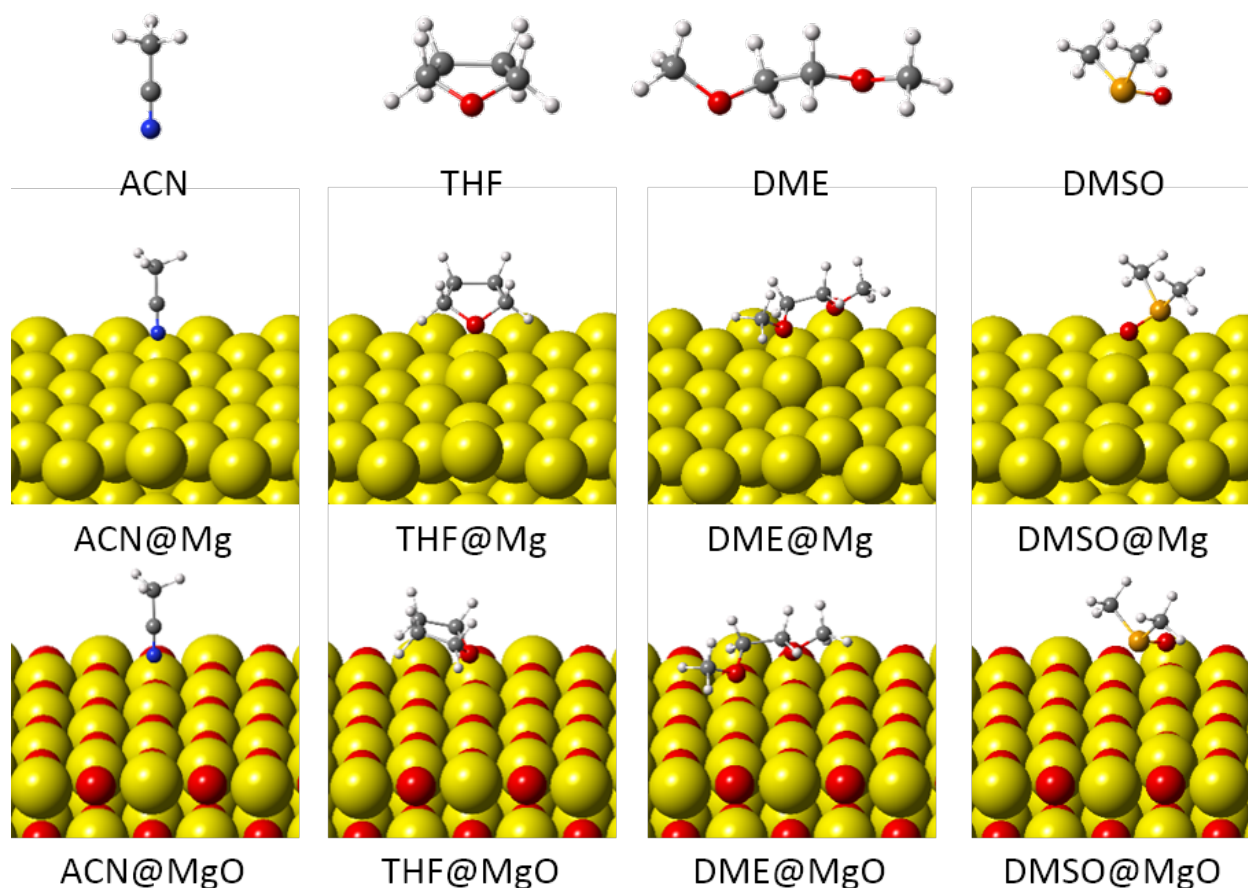


Figure 2. (Top) Relaxed configurations for the isolated solvent molecules examined in this study: ACN, THF, DME and DMSO. (Middle and bottom) Adsorbed configurations for solvent molecules on Mg (0001) and on MgO (001) surfaces. Blue, yellow, orange, grey, red and white spheres represent N, Mg, S, C, O and H, respectively.

Here, quasi-particle techniques based on the GW formalism³⁵⁻³⁶ are used to quantify the impact of interfacial interactions on electrolyte stability for Mg-based batteries. More specifically, we predict the degree of interface-induced renormalization of the HOMO and LUMO levels in a diverse set of four common electrolyte solvents: acetonitrile (ACN), tetrahydrofuran (THF),

dimethoxyethane (DME), and dimethylsulfoxide (DMSO). These solvents span both cyclic and linear geometries (Figure 2, top), while displaying a wide range of dipole moments and dielectric properties: ACN and DMSO exhibit large dielectric constants ($\epsilon_r = 38.8$ and 46.7 , respectively) and dipole moments (both equal to 3.9 D); THF and DME have moderate dielectric constants ($\epsilon_r = 7.6$ and 7.2 , respectively) and small dipole moments (1.6 and 1.7 D, respectively). Mg(0001) and MgO(001) are adopted as model electrode surfaces. The former surface captures the limit of an SEI-free Mg anode, while MgO is used as a proxy for both an oxide-based intercalation cathode and to represent a passivated magnesium anode surface or inorganic SEI layer.

Our calculations indicate that interfacial interactions can alter HOMO/LUMO positions significantly. In molecules exhibiting large dipole moments (ACN and DME), these interactions can reduce the solvent bandgap by up to 25% compared to that of the isolated molecule, signaling a greater tendency for electrochemical decomposition in these cases. Nevertheless, in the absence of solvent-salt interactions, all solvents considered here are predicted to be stable against spontaneous reduction or oxidation by the electrodes. This suggests that these solvents are less susceptible to SEI formation in Mg-based batteries than in an equivalent Li-based cell, where the electrochemical potential in the (Li) anode is more negative. By quantifying the importance of electrolyte/electrode interactions our calculations provide a useful extension to conventional electrolyte screening studies based on isolated molecules.

First-principles calculations were performed using the Vienna *ab initio* Simulation Package (VASP).³⁷⁻³⁸ The projector-augmented wave (PAW) scheme³⁹⁻⁴⁰ was used to treat core–valence electron interactions, with respective valence electronic configurations of $3s^2$ for magnesium, $2s^2 2p^4$ for oxygen, $3s^2 3p^4$ for sulfur, and $2s^2 2p^2$ for carbon. Relaxed configurations for isolated solvent molecules (ACN, THF, DME, and DMSO, Figure 2, top) and for the same molecules

adsorbed on electrode surfaces were calculated using Van der Waals-aware density functionals (vdW-DF2).⁴¹⁻⁴⁴ Electrode surfaces comprised of Mg (0001) and MgO (001) were constructed by laterally-expanding the surface primitive cells 2×3 and 2×2 times, respectively, resulting in surface supercells containing a total of 48 (Mg) and 64 (MgO) atoms. The slab thickness was set at four-layers, with the atom positions in the bottom two layers fixed at their bulk-like positions. All other atoms were allowed to relax until the forces were less than 0.05 eV/Å. Geometry optimization calculations employed a 500 eV energy cutoff for the planewave basis. All surface and interface calculations used a 2×2×1 *k*-point grid. Low-energy adsorption geometries were identified by placing a given solvent molecule at several orientations (vertical, horizontal, and tilted) and at various locations on the surface. The lowest energy configuration was used in subsequent calculations of HOMO/LUMO levels.

The positions of the solvent HOMO and LUMO levels were evaluated using semi-local functionals (PBE-GGA)⁴⁵ and with the many-electron G_0W_0 approximation.^{35-36, 46} In the latter case, wavefunctions from a preceding PBE-GGA calculation were used as input to the non-self-consistent G_0W_0 calculation. Although expensive, the G_0W_0 method yields highly accurate results for bandgaps and band edges.³¹⁻³³ It also incorporates important correlation effects such as image charge interactions that are essential for predicting the energies of HOMO/LUMO levels in molecules adsorbed on metallic surfaces.^{31, 33} These effects are absent in conventional semi-local and hybrid functionals. Our G_0W_0 calculations used approximately 200 empty bands, 32 frequency points and an energy cutoff of 400 eV. We tested these parameters by increasing the number of empty bands to 400, the number of frequency points to 64, and the energy cutoff to 420 eV; no significant shift in the HOMO/LUMO positions resulted from these changes. Consistent positioning of the HOMO/LUMO levels across the different systems was achieved by

aligning the respective vacuum levels of each supercell. Due to the large system sizes investigated, our GW calculations require very large amounts of memory (~ 1.5 TB).

The impact of solvation was examined for the case of DMSO/MgO, using the implicit solvation approach implemented in VASPsol.⁴⁷ Here, self-consistent wavefunctions were re-evaluated at the GGA level in the presence of the continuum solvation field. These wavefunctions were subsequently used as input to G_0W_0 calculations. Only a minor reduction in the electrochemical window (~ 0.1 V) was observed compared to the case where solvation effects were neglected.

Table 1. Calculated adsorption energies (eV/molecule) for solvent molecules on electrode surfaces.

Surface	ACN	THF	DME	DMSO
Mg (0001)	-0.46	-0.61	-0.58	-0.79
MgO (001)	-0.34	-0.67	-0.83	-0.70

The bottom panel of Figure 2 shows the most stable adsorption geometries identified for the four solvent molecules on Mg (0001) and MgO (001). A summary of the calculated adsorption energies is given in Table 1. In all cases the adsorption energy is exothermic with respect to an isolated gas phase molecule. The calculations suggest that ACN has the weakest attraction to the electrode surfaces. It adsorbs with an orientation perpendicular to the surface plane, with its nitrogen atom pointing towards a Mg ion in the surface layer. [The Mg-N separation is 2.21 and 2.34 Å, respectively, for Mg (0001) and MgO (001).] The strongest adsorption interaction occurs for DME on MgO, whereas DMSO exhibits relatively strong bonding with both electrodes. The sulfinyl O of DMSO bonds with surface Mg atoms with a bond length of 2.21 Å on MgO(001) and 2.11 Å on Mg(0001). This distance is similar to the Mg-O distance (2.11 Å) in bulk MgO.

DME and THF both adopt adsorption geometries where the molecule's chain/ring lies parallel to the surface plane.

Table 2. Calculated bandgap (eV) and shift in HOMO/LUMO positions arising from interactions with Mg and MgO surfaces (eV, relative to the vacuum level) in four Mg-electrolyte solvent molecules. Values calculated at the G_0W_0 and GGA levels of theory are reported. For HOMO/LUMO shifts the first value refers to the position of the energy level in the isolated molecule; the second value refers to the (renormalized) position in the adsorbed state. IE and EA are the experimental ionization energy and electron affinity, respectively.

Solvent	Method	Mg (0001)		MgO (001)		HOMO/LUMO gap, E_g (Isolated molecule)	IE	EA
		HOMO	LUMO	HOMO	LUMO			
ACN	G_0W_0	-10.9 \rightarrow -9.7	0.5 \rightarrow -1.1	-10.9 \rightarrow -9.9	0.5 \rightarrow 0.0	11.4	12.2 ⁴⁸	0.0 ⁴⁹
	GGA	-8.0 \rightarrow -9.0	-0.9 \rightarrow -2.6	-8.0 \rightarrow -8.1	-0.9 \rightarrow -1.2	7.1		
THF	G_0W_0	-8.0 \rightarrow -7.6	0.5 \rightarrow -0.1	-8.0 \rightarrow -7.2	0.5 \rightarrow 0.5	8.5	9.4 ⁵⁰	--
	GGA	-5.7 \rightarrow -7.1	-0.9 \rightarrow -1.5	-5.7 \rightarrow -5.8	-0.9 \rightarrow -1.0	4.8		
DME	G_0W_0	-8.1 \rightarrow -7.5	0.2 \rightarrow 0.1	-8.1 \rightarrow -7.3	0.2 \rightarrow 0.5	8.3	9.3 ⁵¹	--
	GGA	-5.8 \rightarrow -7.0	-1.0 \rightarrow -1.4	-5.8 \rightarrow -5.9	-1.0 \rightarrow -0.9	4.8		
DMSO	G_0W_0	-7.3 \rightarrow -7.4	0.4 \rightarrow -1.0	-7.3 \rightarrow -6.8	0.4 \rightarrow 0.2	7.7	9.1 ⁵²	0.0 ⁴⁹
	GGA	-5.3 \rightarrow -6.8	-0.9 \rightarrow -2.4	-5.3 \rightarrow -5.5	-0.9 \rightarrow -1.1	4.4		

Calculations on the isolated molecules were performed in a large cubic box with sufficient vacuum to minimize image interactions. A summary of the calculated HOMO/LUMO gaps is presented in Table 2 for both the GGA and G_0W_0 . The experimental⁴⁸⁻⁵³ ionization energy (IE) and electron affinity (EA), which according to Koopman's theorem approximate the HOMO and LUMO energies,²⁴ respectively, are also reported.

As expected, the semi-local functional severely underpredicts the gaps (by 3-4 eV) in comparison to the G_0W_0 calculations. The largest gap, 11.4 eV as predicted by G_0W_0 , occurs for ACN. This value is within 7% of the experimental gap, estimated as 12.2 eV from the difference IE-EA. Using the coupled electron pair approximation (CEPA⁵⁴), a high-level wavefunction

method, Korth²³ found this gap to be 15.5 eV. With respect to the vacuum level (E_{vac}), the calculated ACN HOMO sits at -10.9 eV, which is slightly more than 1 eV higher than the IE. The G_0W_0 -calculated LUMO is 0.5 eV above both E_{vac} and the EA. In contrast, the gap predicted by the GGA for ACN, 7.1 eV, is more than 4 eV smaller than the G_0W_0 value. In addition, the position of HOMO (LUMO) at the GGA level of theory is ~ 3 (1.5) eV higher (lower) than the corresponding G_0W_0 value. These data clearly indicate the superiority of the G_0W_0 method over semi-local density functionals in reproducing the experimental HOMO/LUMO positions.

For THF and DME the quasi-particle calculations predict similar gaps of 8.5 and 8.3 eV; the positions of their HOMO (LUMO) levels are also similar: -8.0 (0.5) eV for THF and -8.1 (0.2) eV for DME relative to E_{vac} . Although EA data for these molecules appears to be absent from the literature, and therefore a comparison with an experimental value for E_g is not possible, we note that the similarity of the IE values is consistent with the nearly equal values calculated for the HOMO.

Finally, DMSO exhibits the smallest bandgap overall, 7.7 eV based on G_0W_0 , with its HOMO (LUMO) located -7.3 (0.4) eV with respect to E_{vac} . The difference IE-EA suggests a somewhat larger HOMO/LUMO gap of 9.1 eV. In comparison, The CEPA calculations of Korth²³ predicted a gap of 11.5 eV.

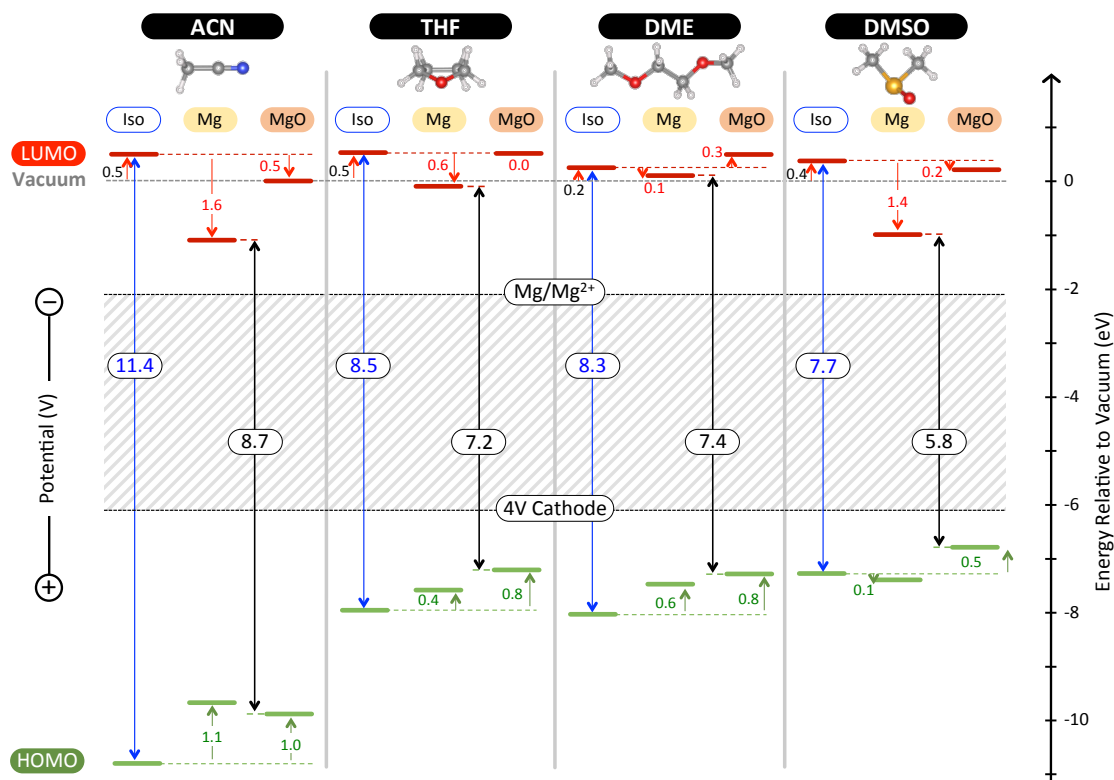


Figure 3. Renormalization of solvent HOMO (green lines) and LUMO (red lines) levels due to interactions with Mg (0001) and MgO (001). The shaded region in the center of the figure represents the electrochemical window of a hypothetical 4V magnesium battery.

A preliminary estimate of the stability of these solvents in an Mg-battery environment can be obtained by comparing the position of the calculated HOMO and LUMO levels to the operating window of a hypothetical Mg cell. This comparison is illustrated in Figure 3, where the HOMO & LUMO levels of the isolated molecules are overlaid with the Mg/Mg^{2+} level of a Mg negative electrode (+2.36 V with respect to SHE) and redox potential of a hypothetical 4 V cathode. It is clear that in the isolated-molecule limit all solvents satisfy the stability criteria illustrated in Figure 1: in all cases $E_g > 4 \text{ eV}$, $E_{\text{LUMO}} > \mu_{\text{A}}$, and $E_{\text{HOMO}} < \mu_{\text{C}}$. The positions of the LUMOs across

all of the molecules vary by only 0.3 eV, suggesting comparable stability against reduction by the anode in the isolated molecule limit.

On the other hand, large differences exist in the HOMO positions, suggesting some variation between molecules in their oxidative stability. Of the solvents examined, ACN is predicted to be the most stable against oxidation due to the deep position (-10.9 eV, Table 2) of its HOMO. In contrast, the HOMO position in the remaining solvents is approximately 3 eV higher, indicative of a higher tendency for oxidation by the cathode. Our calculations suggest that DMSO is the most susceptible to oxidation, whereas THF and DME are marginally more stable.

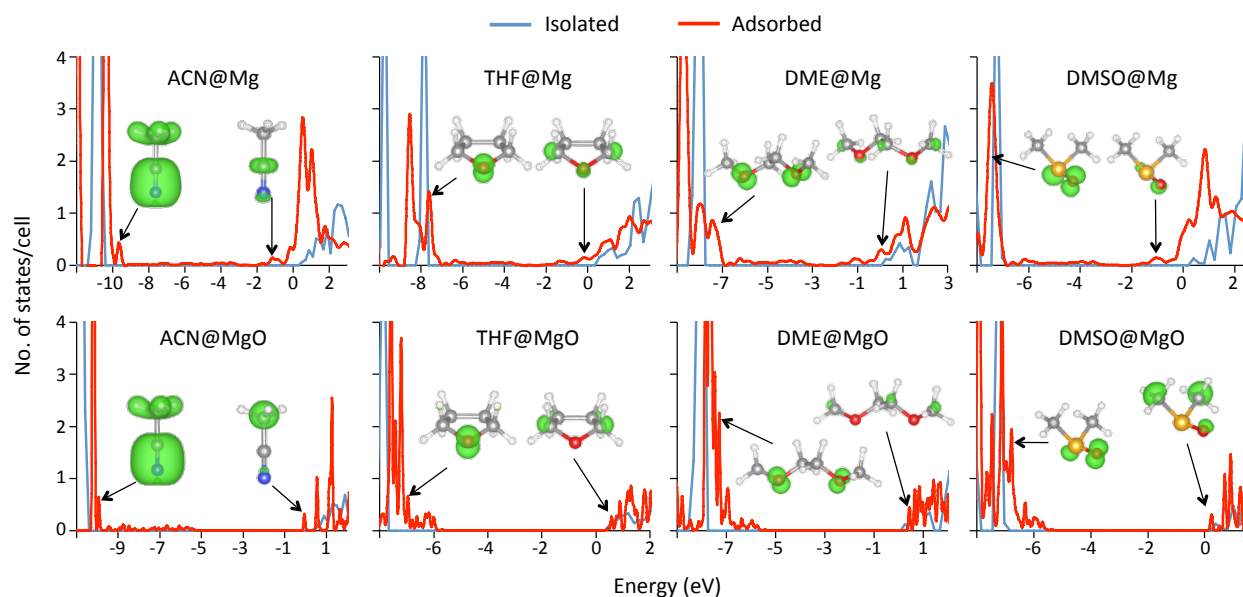


Figure 4. Partial density of states for isolated (blue) and adsorbed (red) solvent molecules on (top) Mg (0001) and (bottom) MgO (001). Insets illustrate the isosurfaces for the HOMO and LUMO for the adsorbed configurations.

How do these conclusions regarding oxidative and reductive stability – which are based on the isolated-molecule limit – hold when interfacial interactions with the electrodes are taken into account? Figure 4 plots the density of states projected onto the solvent molecule in the case of an

isolated molecule (blue curves), and for scenarios depicted in Figure 2, in which the molecule is adsorbed on electrode surfaces represented by Mg (0001) and MgO (001) (red curves). In several cases the solvent's HOMO and LUMO positions shift significantly due to interfacial interactions with the electrode. These shifts are tabulated at the GGA and G_0W_0 levels of theory in Table 2; Figure 3 illustrates the positions of the HOMO and LUMO levels calculated at the G_0W_0 level of theory with respect to the electrochemical window of a hypothetical 4V Mg battery. Compared to the energy levels for the isolated molecules, the trend observed upon adsorption is for the LUMO to shift towards lower energies, while the HOMO shifts to higher energies. Both levels also broaden in the adsorbed state. This general behavior is consistent with earlier GW-based studies of adsorption,^{31,33} and with the broadening of the energy levels of an adsorbate interacting with the broad *sp*-band of the Mg substrate or the oxygen *2p* band of MgO in the weak chemisorption limit.⁵⁵⁻⁵⁹

The observed shifts in the solvent's HOMO and LUMO levels confirm that interactions at the electrolyte/electrode interface can reduce the stability of the electrolyte: interfacial interactions make the electrolyte more susceptible to reduction by the anode and to oxidation by the cathode. In the case of ACN and DMSO, these interactions alter the HOMO-LUMO gap significantly, reducing it by up to 25% of the value for the isolated molecule. A smaller effect is observed for THF and DME, where E_g is reduced by 15% and 11%, respectively. The relatively larger renormalization observed for ACN and THF on Mg can be explained by classical electrostatics. That is, the interaction energy $E(z)$ between a molecule with dipole moment μ and an idealized metal surface is given by $E(z) = -\frac{\mu^2}{8(z-z_{\text{im}})^3}$, where $(z - z_{\text{im}})$ represents the distance from the metal's image plane to the dipole center.⁶⁰ As the strength of interaction scales as μ^2 we expect

that molecules having larger dipole moments (ACN and DMSO) will exhibit a larger degree of renormalization.

We next comment on trends in the reductive and oxidative stability of the individual solvent systems. Interactions with the anode control stability against electrolyte reduction, as the $\text{Mg}^0/\text{Mg}^{2+}$ redox level is much closer to the LUMO levels of the solvents than to their HOMOs. Shifts to the LUMO levels upon adsorption on Mg (0001) and MgO (001) are summarized in the top portion of Fig. 3. Adsorption on the Mg surface captures the limit of a metallic anode, whereas the MgO surface serves as a proxy for an electronically insulating passivation layer (or SEI) on an underlying Mg anode. In general, interactions with either surface tend to lower the LUMO level, however renormalization of the LUMO is largest for interactions with metallic Mg. In this case the calculated LUMO shifts are particularly large for ACN (1.6 eV) and DMSO (1.4 eV), as previously described. In the former case the renormalized LUMO is within ~ 1 V of the $\text{Mg}^0/\text{Mg}^{2+}$ level, suggesting that ACN is the most likely solvent to undergo reductive decomposition out of the four systems considered. The relatively large downward shifts predicted for the LUMO levels in ACN and DMSO upon adsorption on Mg(0001) are consistent with experiments suggesting that these solvents are reduced on magnesium electrodes at relatively small negative voltages (*vs.* Mg/Mg^{2+}).^{61,62}

In contrast to the relatively large shifts on Mg, renormalization of the LUMO levels on MgO are much smaller, with a maximum reduction of 0.5 eV occurring for ACN. The smaller shift on the oxide is expected due to the much smaller contribution of image charge interactions in insulating surfaces such as MgO.^{8, 60} We conclude that the presence of an insulating film on the Mg anode will improve the reductive stability of the electrolyte. However, as previously

mentioned, the presence of such a film may have other undesirable side effects, such as limited conductivity for Mg ions.¹⁹⁻²⁰

Interactions with the battery's cathode control the oxidative stability of the electrolyte. This is due to the fact that the HOMO position is much closer to the redox level of the cathode, μ_C , than to the $\text{Mg}^0/\text{Mg}^{2+}$ level. As previously described, MgO (100) is adopted as a model for the surface of an oxide-based cathode. For oxidative stability the HOMO should sit at a more positive potential (i.e., lower in energy) than the cathode redox potential. This requirement is satisfied by all solvents considered, Fig. 3. However the degree by which this requirement is met varies from solvent to solvent. Based on its behavior in isolation, ACN is the most stable solvent against oxidation. Nevertheless, its HOMO is renormalized the most (i.e., is shifted 1 eV higher) in the presence of MgO. THF and DME display similar HOMO shifts of 0.8 eV, resulting in similar oxidative stability. Finally, despite having the smallest interface-induced shift to its HOMO (0.5 eV), DMSO exhibits the worst oxidative stability overall, as it is only 0.7 V below μ_C . This is largely a consequence of the high position of its HOMO in the isolated molecule limit.

Although interfacial interactions clearly reduce stability, the absolute positions of the HOMO & LUMO levels with respect to the electrochemical window of the Mg-battery (Fig. 3) suggest that all four solvents remain stable in an *absolute* sense. That is, even with interface-induced level shifts, the stability criteria remain satisfied: $E_g > 4$ eV, $E_{\text{LUMO}} > \mu_A$, and $E_{\text{HOMO}} < \mu_C$. As previously mentioned, these criteria are not always satisfied in Li-ion batteries, resulting in the formation of an SEI that kinetically stabilizes the electrolyte/electrode interface. The data presented here suggest that an SEI *should not* form for a Mg-based system, at least for the scenario considered here. (Our model considers only solvent/electrode interactions, whereas a realistic electrolytic solution also contains salt molecules. The salt can interact with the solvent,

and can potentially contribute to SEI formation on its own through redox reactions with the electrode. In addition, heterogeneities on the electrode surface such as vacancies or steps could also serve as sites having a higher activity for electrolyte decomposition. In principle, all of the above factors should be accounted for in a realistic description of electrolyte stability in a functioning electrochemical cell.) Quantifying solvent/electrode interactions, as is done here, is a step toward achieving this understanding.

To summarize, first-principles many-body perturbation theory has been used to predict changes in the electrochemical window of a diverse set of common electrolyte solvents arising from interfacial interactions with Mg-based electrodes. Understanding these interfaces is critical to maximizing battery performance, as they are the location where the oxidation or reduction of electrolyte species will occur. Nevertheless, relatively little is known about the degree renormalization of HOMO/LUMO levels at electrochemical interfaces; due to their complexity, these effects are typically overlooked in studies that screen for optimal electrolytes.

The systems examined span a wide range of geometries and electronic properties. The solvents (THF, DME, DMSO and ACN) include both cyclic and linear geometries, and exhibit a diverse sampling of dipole moments and dielectric constants. Electrode surfaces include metallic Mg and insulating MgO; the former captures the limit of an unpassivated negative electrode, while the latter serves as a model for a passivated Mg surface (or an inorganic SEI), or for an oxide-based positive electrode.

Although computationally demanding, the GW-based methods employed here represent the state-of-the-art for predicting energy levels in condensed-phase systems. These methods dramatically out-perform conventional semi-local functionals, and are shown to be in good agreement with experimental measurements of ionization potentials and electron affinities.

Our calculations indicate that the degree of narrowing or ‘renormalization’ of the electrochemical window of solvent molecules at interfaces can be substantial. For example, interface effects reduce the HOMO-LUMO gap by approximately 25% in cases involving molecules with large dipole moments (ACN & DMSO) adsorbed on Mg. This behavior is consistent with the greater tendency for reductive decomposition of ACN and DMSO on Mg anodes. In contrast, the degree of renormalization is smaller for interfaces involving MgO and/or solvents exhibiting small dipoles. This trend is roughly consistent with a simple electrostatic description of a dipole interacting with a solid surface.

These renormalizations signal that interfacial interactions promote redox of these molecules. Nevertheless, neglecting interactions with the electrolyte salt, comparisons with the electrochemical potentials of a hypothetical 4V Mg-battery indicate that these solvents are stable in an *absolute* sense, i.e., $E_g > 4V$, $E_{LUMO} > \mu_A$, and $E_{HOMO} < \mu_C$. We conclude that the driving force for SEI formation (due to electrolyte decomposition) in a Mg-battery should be smaller than in an equivalent Li-battery. This would appear to be welcome news, as Aurbach has argued that Mg-based SEIs are most likely poor conductors of Mg-ions.

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The authors declare no competing financial interests.

ACKNOWLEDGMENT

This work was supported as part of the Joint Center for Energy Storage Research, an Energy Innovation Hub funded by the U.S. Department of Energy, Office of Science, Basic Energy Sciences. D.J.S. acknowledges DTU Energy, the Villum Foundation's Visiting Professor Program, and the Nordea Foundation's Residence Program for support during his stay at DTU.

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